NDC NUMBER FOR FLULAVAL CPT 90688 Latest Edition

ndc number for flulaval cpt 90688 is a crucial detail for healthcare providers, insurance companies, and medical billing specialists involved in the administration and reimbursement of influenza vaccinations. Understanding the National Drug Code (NDC) number associated with Flulaval, especially when administered using CPT code 90688, is essential for accurate billing, compliance, and seamless processing of insurance claims. This article offers a comprehensive overview of the NDC number for Flulaval CPT 90688, its significance in medical billing, how to identify the correct NDC, and related information to optimize your knowledge and practice.

Understanding the NDC Number and CPT Code 90688

What is an NDC Number?

The National Drug Code (NDC) is a unique 10-11 digit identifier assigned to medications, including vaccines, by the Food and Drug Administration (FDA). It helps to accurately identify drug products, their manufacturer, and specific packaging details. The NDC number facilitates tracking, billing, inventory management, and regulatory compliance.

Key Points about NDC Numbers:

- Comprise three segments: Labeler Code, Product Code, and Package Code
- Standardized format ensures consistency across healthcare systems
- Used in pharmacy and medical billing to specify exact drug formulations and packaging

What is CPT Code 90688?

CPT (Current Procedural Terminology) code 90688 is used to report the administration of the quadrivalent influenza vaccine, Flulaval, for patients aged 3 years and older. This code covers the injection procedure and is essential for billing insurance providers for vaccine administration.

Highlights of CPT 90688:

- Indicates the administration of Flulaval quadrivalent vaccine
- Applicable for patients 3 years and older

- Often bundled with the vaccine product itself for billing purposes

The Significance of the NDC Number for Flulaval CPT 90688

Why is the NDC Number Important?

Having the correct NDC number linked with CPT 90688 ensures that billing is precise and reimbursement is smooth. It confirms the exact vaccine product used, which is critical for:

- Verifying vaccine authenticity
- Ensuring coverage and compliance with healthcare policies
- Avoiding claim denials or rejections
- Tracking vaccine distribution and usage

How the NDC Number Interacts with CPT Codes

While CPT codes describe the procedure (e.g., vaccine administration), the NDC specifies the exact vaccine product administered. Insurance payers often require the NDC number on claims to verify the appropriateness of the billed vaccine.

Example Workflow:

- Administer Flulaval quadrivalent vaccine (CPT 90688)
- Document the specific NDC number of the vaccine used
- Submit claim with both the CPT code and NDC number
- Insurance processes the claim based on both identifiers

Locating the NDC Number for Flulaval

Where to Find the NDC Number

The NDC number can be located on:

- The vaccine packaging
- The manufacturer's label or package insert
- Electronic health records or billing software

Steps to identify the correct NDC:

- Check the vaccine box or vial label for the NDC number
- Confirm the NDC matches the specific lot and packaging used
- Record the NDC accurately in the medical record and billing documentation

Examples of NDC Numbers for Flulaval

Although NDC numbers may vary based on manufacturer packaging, some examples include:

- 58160-0770-30
- 58160-0770-02
- 58160-0770-60

Note: Always verify with the current manufacturer?s packaging or official sources for the most accurate and up-to-date NDC.

How to Use the NDC Number for Billing and Reimbursement

Steps for Proper Billing

To ensure accurate billing for Flulaval administered with CPT 90688:

- Record the patient's details and the date of vaccination
- Document the specific NDC number of the vaccine used
- Use the CPT 90688 code for the administration procedure
- Submit the claim including both CPT and NDC codes
- Attach any required documentation or supporting information

Common Challenges and Solutions

- Incorrect NDC Entry: Double-check the NDC before submitting claims
- Outdated NDC Codes: Ensure using the latest NDC numbers from the manufacturer
- Reimbursement Issues: Confirm payer requirements for NDC submission and follow their guidelines

Benefits of Using the Correct NDC Number for Flulaval

Advantages include:

- Improved claim approval rates
- Accurate inventory and supply chain management
- Better tracking of vaccine usage and outcomes
- Compliance with federal and state reporting requirements

Additional Resources and Best Practices

Where to Find Updated NDC Data

- FDA?s NDC Directory
- Manufacturer websites
- Electronic health record (EHR) systems
- Medical billing software with integrated NDC databases

Best Practices for Healthcare Providers

- Always verify NDC numbers at the point of vaccination
- Keep updated NDC lists from manufacturers
- Educate billing staff on the importance of NDC accuracy
- Use electronic billing systems that support NDC entry

Conclusion

The NDC number for Flulaval CPT 90688 is more than just a set of digits; it is a vital component of accurate healthcare documentation, billing, and compliance. Properly identifying and recording the correct NDC ensures that providers receive appropriate reimbursement, maintains compliance with regulations, and supports the effective management of vaccine inventories. As influenza vaccines continue to be a cornerstone of public health efforts, understanding the relationship between NDC numbers and CPT codes like 90688 is essential for all healthcare professionals involved in vaccine administration and billing.

By consistently verifying and using the correct NDC numbers, healthcare providers can streamline their billing processes, improve patient care documentation, and contribute to better health outcomes. Whether you are a clinician, billing specialist, or administrator, staying informed about the NDC number for Flulaval and its connection to CPT 90688 is a fundamental aspect of modern healthcare practice.

NDC Number for Flulaval CPT 90688: A Comprehensive Guide

Understanding the intricacies of healthcare billing and coding is essential for providers, insurers, and patients alike. One key element in this process is the National Drug Code (NDC) number, a unique identifier for medications, which ensures precise communication across the healthcare system. When it comes to the flu vaccine administered under CPT code 90688, knowing the correct NDC number for Flulaval is crucial for accurate billing, reimbursement, and inventory management. In this guide, we will explore the significance of the NDC number for Flulaval CPT 90688, how to locate it, and best practices for its use in clinical and billing settings.

What is the NDC Number and Why is it Important?

Definition and Purpose of NDC Numbers

The National Drug Code (NDC) is a standardized 10- or 11-digit number assigned by the Food and Drug Administration (FDA) to identify medications, including vaccines, in the United States. It serves as a universal product identifier, providing details about the drug?s manufacturer, product, and packaging.

Significance in Healthcare

- Billing and Reimbursement: Accurate NDCs ensure that insurance claims are processed smoothly, reducing rejections due to mismatched or missing data.
- Inventory Management: Pharmacies and healthcare providers rely on NDCs to track stock levels and manage supplies effectively.
- Regulatory Compliance: Proper coding helps meet federal and state reporting requirements, especially for vaccination data.

Understanding CPT Code 90688 and Its Association with Flulaval

What Does CPT 90688 Cover?

CPT code 90688 refers specifically to the Influenza virus vaccine, quadrivalent (IIV4), split virus, preservative free, preservative added, administered via intramuscular injection. It is commonly used during flu season to bill for flu vaccinations.

Connection with Flulaval

Flulaval is a brand of quadrivalent influenza vaccine, often administered under CPT 90688. To bill for a Flulaval vaccination correctly, providers must supply the corresponding NDC number that matches this specific vaccine.

The NDC Number for Flulaval: How to Find It

Factors Influencing the NDC Selection

- Vaccine Formulation: Flulaval comes in various formulations, including dosage strength and packaging.
- Manufacturer Variations: The NDCs vary depending on the manufacturing batch and packaging size.
- Updated Listings: NDCs are periodically updated; always verify the latest codes.

Sources to Locate the Correct NDC for Flulaval

- FDA?s NDC Database: The official source for all approved NDCs.
- Vaccine Packaging: The NDC is printed on the vaccine box or vial label.
- Pharmaceutical Manufacturer?s Documentation: The manufacturer's product inserts or websites.
- Healthcare Billing Resources: Reputable coding and billing guides often list NDCs associated with CPT codes.

Example of NDCs for Flulaval

While the exact NDC can vary, typical NDCs for Flulaval quadrivalent influenza vaccine might include:

- 49281-0520-xx
- 49281-0521-xx

(Note: "xx" indicates different packaging sizes, e.g., 10-dose vials, pre-filled syringes, etc.)

Always confirm the specific NDC matching your product batch and packaging.

Best Practices for Using NDC Numbers with CPT 90688

Accurate Documentation

- Record the Correct NDC: Ensure the chosen NDC matches the vaccine administered.
- Include NDC in Billing Claims: Many payers require the NDC to be listed on the claim form (e.g.,

CMS-1500 or UB-04).

Verification Before Administration

- Check the NDC at the Point of Care: Confirm the NDC on the vaccine packaging before administration.
- Update Electronic Medical Records (EMRs): Enter the correct NDC to facilitate accurate billing and reporting.

Stay Updated

- Regularly Review NDC Listings: Because codes can change, consult current resources annually or as needed.
- Coordinate with Suppliers: Ensure your vaccine suppliers provide the correct NDC information.

Why Proper NDC Usage Matters

Ensuring Reimbursement

Insurance companies often require an exact match between the NDC and the CPT code to process claims successfully. Using an incorrect or outdated NDC can lead to claim denials or delays.

Facilitating Data Reporting

Accurate NDCs support public health reporting, vaccine tracking, and inventory management, which are critical during flu seasons and vaccination campaigns.

Avoiding Compliance Issues

Misrepresentation or incorrect coding can lead to compliance violations, audits, or penalties.

Practical Tips for Providers and Billing Professionals

- Create a Reference List: Maintain an up-to-date list of NDCs associated with common vaccines like Flulaval.
- Use EHR/EMR Integration: Leverage electronic systems that prompt for NDC entry during vaccine administration.
- Train Staff: Educate clinical and billing staff on the importance of matching NDCs with CPT

codes.

- Audit Regularly: Conduct periodic reviews of vaccination records and billing submissions to catch discrepancies early.

Conclusion

The NDC number for Flulaval CPT 90688 is a vital component in the complex process of healthcare billing and inventory management. Correctly identifying and documenting the NDC ensures smooth reimbursement, regulatory compliance, and accurate public health reporting. Healthcare providers must stay informed about the latest NDC listings, verify the specific code that aligns with their vaccine supply, and incorporate these details into their clinical and administrative workflows. By doing so, they can streamline operations, reduce billing errors, and contribute to more effective vaccination programs.

Remember: Always consult current FDA resources, manufacturer documentation, and trusted coding guides to ensure your NDC information remains accurate and up-to-date.

QuestionAnswer What is the NDC number for Flulaval in relation to CPT code 90688? The NDC number for Flulaval associated with CPT 90688 varies by manufacturer and packaging, but it typically can be found on the vaccine packaging or by consulting the manufacturer's product information. Why is the NDC number important for billing Flulaval with CPT 90688? The NDC number uniquely identifies the vaccine product, ensuring accurate billing and reimbursement when submitting claims with CPT 90688 for influenza vaccination. Where can I find the NDC number for Flulaval to use with CPT code 90688? The NDC number is usually printed on the vaccine packaging or label. It can also be obtained from the vaccine manufacturer's website or the FDA's NDC database. Is the NDC number the same for all Flulaval vaccines used with CPT 90688? No, different formulations or packaging sizes of Flulaval may have different NDC numbers. Always verify the specific NDC for the product being administered. Can I use any NDC number for Flulaval when billing CPT 90688? No, you should use the specific NDC number that corresponds to the exact vaccine product and packaging used during administration to ensure proper billing. Does CMS require the NDC number when billing CPT 90688 for Flulaval vaccinations? Yes, CMS generally requires the NDC number to be submitted with claims for vaccines like Flulaval to ensure proper identification and reimbursement. Are there updates to the NDC numbers for Flulaval associated with CPT 90688? NDC numbers can be updated or changed by manufacturers; it's important to verify the current NDC at the time of billing, often through the vaccine manufacturer or official databases. How can I verify the correct NDC number for billing Flulaval with CPT 90688? You can verify the correct NDC by consulting the vaccine manufacturer, reviewing the packaging label, or checking the FDA's National Drug Code database for the most accurate and current information.

Related keywords: NDC number, Flulaval, CPT code 90688, influenza vaccine, vaccination coding, NDC lookup, flu shot billing, immunization codes, vaccine reimbursement, vaccine inventory

chemistry worksheet atomic number and mass number

Chemistry worksheet atomic number and mass number are fundamental concepts in the study of chemistry that help students and enthusiasts understand the structure of atoms. These concepts are essential for grasping how elements are organized, how isotopes differ from each other, and how atomic properties influence chemical behavior. A well-designed chemistry worksheet focusing on atomic number and mass number serves as an excellent educational tool to reinforce these foundational ideas, making complex topics more accessible through practical exercises, clear explanations, and engaging activities.

Understanding Atomic Number and Mass Number

What is the Atomic Number?

The atomic number (denoted as Z) is a fundamental property of an element. It represents the number of protons present in the nucleus of an atom of that element. Since protons are positively charged particles, the atomic number essentially defines the element itself because each element has a unique number of protons.

- Determines the element's identity (e.g., Hydrogen has Z=1, Carbon Z=6).
- Impacts the element?s chemical properties and reactivity.
- Defines the number of electrons in a neutral atom.

What is the Mass Number?

The mass number (denoted as A) refers to the total number of protons and neutrons in an atom's nucleus. Unlike the atomic number, which is unique for each element, the mass number can vary among isotopes of the same element.

- Calculates the approximate atomic mass of an isotope.
- Used to distinguish between different isotopes.
- Not a whole number for elements with naturally occurring isotopic mixtures, where average atomic mass is used instead.

Key Differences Between Atomic Number and Mass Number

FeatureAtomic Number (Z)Mass Number (A)
DefinitionNumber of protons in an atomTotal number of protons and neutronsSymbolZAUnique for

each elementYesNo (varies among isotopes)DeterminesElement identityIsotope type

Using a Chemistry Worksheet to Learn Atomic Structure

A chemistry worksheet centered on atomic number and mass number provides structured exercises that enhance understanding through various activities such as:

- Identifying elements based on atomic number
- Calculating the number of neutrons in isotopes
- Determining the isotope symbol notation
- Matching isotopes to their atomic and mass numbers
- Understanding the concept of atomic mass as an average of isotopes

These activities help students master key concepts and develop problem-solving skills essential for advanced chemistry topics.

How to Calculate Atomic Number and Mass Number

Calculating Atomic Number

The atomic number is usually given or can be found on the periodic table. When not provided directly, it is often determined by:

- Looking at the element symbol (e.g., C for Carbon, atomic number 6)
- Using the periodic table
- Counting the number of protons in an atom if data is provided

Calculating Mass Number

The mass number can be calculated if the number of protons and neutrons are known:

Mass Number (A) = Number of Protons + Number of Neutrons

Alternatively, for a specific isotope, the notation is written as:

Element symbol with mass number: ^{12}C (Carbon-12)

where 12 indicates the mass number.

Isotopes and Their Significance in Chemistry

What are Isotopes?

Isotopes are variants of a particular chemical element that have the same number of protons but different numbers of neutrons. This variation results in different mass numbers but similar chemical properties.

- Examples: Carbon-12 (^{12}C), Carbon-13 (^{13}C), and Carbon-14 (^{1}?C).
- Some isotopes are stable, while others are radioactive.

Importance of Isotopes

Isotopes have practical applications in various fields:

- Radioactive dating (e.g., Carbon-14 dating in archaeology)
- Medical imaging and treatment (e.g., Radioisotopes in cancer therapy)
- Tracing chemical reactions in research
- Understanding atomic structure and nuclear physics

Practice Exercises for Students

To reinforce learning, a chemistry worksheet should include practice questions such as:

- Identify the atomic number and mass number of an element given its isotope notation (e.g., ??Ca).
- Calculate the number of neutrons in the isotope ?Li, given that Lithium?s atomic number is 3.
- Write the isotope notation for an atom with 11 protons and 12 neutrons.
- Determine whether a given isotope is stable or radioactive based on its atomic and mass numbers.
- Compare two isotopes of the same element and discuss how their properties differ.

Answers to these questions help students develop confidence and deepen their understanding of atomic concepts.

Visual Aids and Diagrams in the Worksheet

Incorporating visual elements enhances comprehension:

- Diagrams of atomic structure highlighting protons, neutrons, and electrons.
- Periodic table with atomic numbers and isotopic information.
- Charts showing common isotopes of key elements.
- Visual comparisons between isotopes to illustrate differences in neutrons.

Such visuals make abstract concepts more tangible and facilitate better retention.

Additional Resources and Study Tips

Students seeking to excel in understanding atomic number and mass number should consider:

- Using periodic tables effectively, noting atomic numbers and isotopic masses.
- Practicing calculations with real data from elements and isotopes.
- Engaging in hands-on activities, such as building atom models.
- Utilizing online interactive quizzes and simulations for practice.

Consistent practice with a variety of problems enhances mastery and prepares students for higher-level chemistry topics.

Conclusion

A comprehensive chemistry worksheet on atomic number and mass number is an invaluable educational resource that reinforces core atomic concepts vital for understanding the nature of matter. By exploring definitions, differences, calculations, and practical applications, learners gain a solid foundation in atomic structure. Integrating visual aids, practice exercises, and real-world examples within the worksheet makes learning engaging and effective. Mastery of these concepts not only prepares students for exams but also deepens their appreciation of the intricate building blocks of the universe, paving the way for success in advanced chemistry and related sciences.

Chemistry Worksheet Atomic Number and Mass Number: An In-Depth Exploration

In the realm of chemistry, understanding the fundamental building blocks of matter is essential for grasping how elements interact, form compounds, and underpin the universe itself. Among the foundational concepts are the atomic number and mass number, which serve as key identifiers and descriptors of atoms. These numerical values not only facilitate the classification of elements but also underpin various calculations and models within atomic physics and chemistry. This article offers a comprehensive review of these concepts, exploring their definitions, significance, differences, and applications, while providing an analytical perspective on their roles in chemical education and research.

Understanding Atomic Number and Mass Number: Definitions and Significance

What is the Atomic Number?

The atomic number, symbolized as Z, is a fundamental property of an element that specifies the number of protons in the nucleus of an atom. Since protons carry a positive charge, the atomic number directly determines the element's identity. For example, all atoms with an atomic number of 6 are carbon atoms, while those with an atomic number of 8 are oxygen atoms.

Significance of the Atomic Number:

- Element Identification: The atomic number uniquely identifies an element; no two different elements share the same atomic number.
- Periodic Table Organization: Elements are arranged in order of increasing atomic number, which structures the periodic table and reflects periodic trends.
- Chemical Properties: The atomic number influences an atom?s electron configuration, dictating its chemical reactivity, bonding behavior, and valency.

What is the Mass Number?

The mass number, denoted as A, represents the total number of nucleons?protons and neutrons?in an atom's nucleus. Unlike the atomic number, which is fixed for each element, the mass number can vary within the same element, leading to different isotopes.

Significance of the Mass Number:

- Isotope Identification: Variations in the mass number give rise to isotopes, which are atoms of the same element with different neutron counts.
- Mass Calculations: The mass number is used to approximate the atomic mass of an isotope, which is vital for calculations in stoichiometry and nuclear chemistry.
- Radioactivity and Stability: The ratio of neutrons to protons (reflected in the mass number) influences nuclear stability and the likelihood of radioactive decay.

Distinguishing Between Atomic Number and Mass Number: Key Differences

Understanding the differences between atomic number and mass number is crucial for students and professionals alike. Below are the primary distinctions:

| Aspect | Atomic Number (Z) | Mass Number (A) |

|---|---|---|

| Definition | Number of protons in the nucleus | Total number of protons and neutrons in the nucleus |

| Units | No units; a count | No units; a count |

| Fixed for an element | Yes | No; varies with isotopes |

| Determines | Element identity | Isotope identity |

| Affects | Chemical properties | Nuclear properties and atomic mass |

Additional insights:

- Since the atomic number determines the element, changing Z results in a different element altogether.
- The mass number can vary within the same element, leading to different isotopes with distinct physical properties.

Calculating Atomic Mass and Isotopic Composition

The Role of Isotopes in Atomic Mass Calculations

While the atomic mass listed on the periodic table is a weighted average, individual isotopes have specific masses approximately equal to their mass numbers. Understanding isotopic composition is essential for applications in:

- Radiometric dating
- Medical imaging
- Nuclear energy

Calculating Atomic Mass:

The atomic mass of an element is calculated as the weighted average of the masses of its naturally occurring isotopes:

$$

\text{Atomic mass} = \sum (\text{fractional abundance of isotope} \times \text{mass of isotope})

$$

For example, chlorine has two main isotopes:

- Chlorine-35 (mass number 35, abundance ~75%)
- Chlorine-37 (mass number 37, abundance ~25%)

The average atomic mass:

$$

(0.75 \times 35) + (0.25 \times 37) = 26.25 + 9.25 = 35.5 \text{ amu}

$$

This explains why chlorine?s atomic mass appears as approximately 35.5 amu on the periodic table.

Implications for Chemical and Physical Properties

Isotopic differences influence physical properties such as:

- Atomic weight
- Spectral lines
- Nuclear stability

In chemical reactions, isotopic substitution can lead to subtle effects, such as isotope effects, affecting reaction rates and mechanisms.

Using Atomic Number and Mass Number in Practice

Notation and Representation

Atoms are commonly represented using nuclear notation:

$$

^{A}_{Z}\text{X}

\]

Where:

- X is the chemical symbol
- Z is the atomic number
- A is the mass number

For example, an isotope of carbon with 6 protons and 7 neutrons is written as:

\[

^{13}_{6}\text{C}

\]

This notation succinctly conveys the isotope's identity and composition.

Determining Neutron Number

Given the atomic number and mass number, the neutron count can be calculated:

\[

\text{Number of neutrons} = A - Z

\]

Using the previous example:

\[

7 = 13 - 6

\]

Understanding neutron numbers is critical in analyzing nuclear stability and decay processes.

Applications in Radioactive Decay and Nuclear Reactions

- Radioactive isotopes are characterized by unstable nuclei with specific neutron-to-proton ratios.
- Nuclear reactions often involve changes in the mass and atomic numbers, such as in nuclear fission or fusion.
- Detection and measurement of isotopes rely on their unique atomic and mass numbers.

Educational and Practical Importance of the Worksheet Approach

Using worksheets to teach atomic number and mass number enhances comprehension through active engagement. They often include exercises such as:

- Filling in atomic and mass numbers for different isotopes
- Calculating neutrons and protons
- Converting between notation forms
- Interpreting atomic data for element identification

Benefits of such worksheets include:

- Reinforcing theoretical knowledge
- Developing problem-solving skills
- Preparing students for laboratory and exam scenarios
- Facilitating deeper understanding of atomic structure

Advanced Topics and Contemporary Research

Beyond foundational concepts, modern research in nuclear physics and chemistry explores:

- Superheavy elements: Atoms with extremely high atomic and mass numbers, often synthesized in laboratories.
- Isotope enrichment: Techniques to increase the proportion of specific isotopes for medical or industrial use.
- Nuclear medicine: Utilization of radioactive isotopes (e.g., technetium-99m) for diagnostics and treatment.
- Nuclear stability theories: Models predicting the stability and decay pathways based on neutron-proton ratios.

These advanced topics demonstrate how the understanding of atomic and mass numbers continues to evolve, impacting technology and medicine.

Conclusion: The Integral Role of Atomic and Mass Numbers in Chemistry

The concepts of atomic number and mass number are cornerstones of atomic theory and chemistry. They serve as precise identifiers for elements and isotopes, underpin calculations in stoichiometry, and inform our understanding of nuclear phenomena. As educational tools, worksheets that focus on these numbers cultivate critical thinking and foundational knowledge, essential for any aspiring chemist, physicist, or researcher. Recognizing their distinctions and applications enhances our comprehension of matter at its most fundamental level and underscores the intricate beauty of the atomic world.

In summary, mastery of atomic and mass numbers is not merely academic but vital for advancing scientific inquiry, technological innovation, and practical applications across multiple disciplines.

QuestionAnswer What is the atomic number of an element? The atomic number of an element is the number of protons found in the nucleus of an atom of that element. How is the mass number of an atom different from its atomic number? The mass number is the total number of protons and neutrons in an atom's nucleus, whereas the atomic number only counts the protons. Why is the atomic number important in identifying elements? Because each element has a unique atomic number, it serves as the primary identifier for distinguishing one element from another. Can the mass number of an atom change? Yes, the mass number can change due to the addition or removal of neutrons, resulting in different isotopes of the same element. How do you calculate the atomic number and mass number from an element's symbol? The atomic number is usually written as a subscript before the symbol, and the mass number as a superscript. For example, in ?^{2}H, 1 is the atomic number and 2 is the mass number. What is an isotope, and how is it related to atomic and mass numbers? An isotope is an atom of the same element with the same atomic number but a different mass number due to varying neutrons. Why are isotopes of an element chemically similar? Because isotopes have the same number of protons and electrons, they exhibit similar chemical behavior despite differences in mass. How can understanding atomic and mass numbers help in scientific studies? It helps in identifying elements and their isotopes, understanding nuclear reactions, and calculating atomic weights for chemical calculations.

Related keywords: atomic number, mass number, isotopes, protons, neutrons, electrons, atomic structure, periodic table, atomic mass, element symbols

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